

## A Single-Crystal Neutron-Diffraction Study of Urea-Phosphoric Acid\*

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The 1:1 compound of urea and phosphoric acid,  $(\text{NH}_2)_2\text{CO}\cdot\text{H}_3\text{PO}_4$ , has been studied by single-crystal neutron-diffraction techniques. All atoms, including the hydrogen atoms, have been accurately located. The carbonyl oxygen atom of urea is linked to the phosphoric acid moiety by a short hydrogen bond in which an acid hydrogen atom is nearly centered. Each of the remaining hydrogen atoms participates in an ordinary hydrogen bond. The space group is orthorhombic *Pbca* with lattice parameters  $a=17\cdot6492$  (6),  $b=7\cdot4820$  (3), and  $c=9\cdot0018$  (4) Å at 23°C.

### Introduction

Urea and phosphoric acid form a 1:1 crystalline compound,  $(\text{NH}_2)_2\text{CO}\cdot\text{H}_3\text{PO}_4$ , which has been examined by X-ray diffraction techniques by Sundera-Rao, Turley & Pepinsky (1957) and independently by Wolfram, Arutunian, Antishkina, & Porai-Koshits (1967). The symmetry is that of the orthorhombic space group *Pbca*, and there are eight formula units per cell.

Although the early studies did not reveal the hydrogen atom positions, they did show a short O...O distance between the two moieties. Sundera-Rao, Turley and Pepinsky suggested that this represents a strong hydrogen bond in which the hydrogen atom is nearly equidistant from the two oxygen atoms.

We report here a single-crystal neutron-diffraction study of this compound which accurately locates all the hydrogen atoms and clearly shows this centered hydrogen bond. After our work was completed we learned that Mootz & Albrand (1972) had made a precise X-ray diffraction study of this material which also located the hydrogen atoms. There are no major discrepancies between their results and ours.

Two other urea-acid compounds which have been studied recently also have hydrogen bonds between the acid and the carbonyl oxygen of urea. In the 1:1 compound formed from urea and nitric acid, Worsham & Busing (1969) found that the acid proton had transferred completely from the nitrate group to the carbonyl oxygen. This material should therefore be classed as a salt, uronium nitrate.

In the 2:1 compound of urea and oxalic acid (Harkema, Bats, Weyenberg & Feil, 1972) on the other hand, the acid proton remains as part of the oxalic acid molecule and forms an ordinary hydrogen bond. This material is therefore an addition compound.

The urea-phosphoric acid of the present study must be regarded as intermediate between these two ex-

tremes. It is clear that the differences between the structures of these compounds must be related qualitatively to the relative strengths of the acids involved, but we are not prepared to present a quantitative theory at this time.

### Experimental

Sample crystals were obtained by slow evaporation of an aqueous solution containing equimolar quantities of urea and phosphoric acid. Lattice parameters were redetermined from X-ray diffractometer measurements of 12 high-angle reflections using the method of least-squares (Busing & Levy, 1967). At 23°C they are  $a=17\cdot6492$ (6)  $b=7\cdot4820$  (3), and  $c=9\cdot0018$  (4) Å, based on a wavelength of 1.54051 Å for Cu  $K\alpha_1$  radiation.

The sample for neutron diffraction weighed 69.3 mg and measured between 3 and 5 mm in each dimension. Integrated intensities of 1926 independent reflections were measured on the Oak Ridge automatic diffractometer (Busing, Smith, Peterson & Levy, 1964). These were put on an absolute scale by comparison with the intensity from a NaCl standard crystal of known weight. The temperature of measurement was 25°C, and the neutron wavelength was 1.077 Å.

Transmission factors which ranged from 0.45 to 0.59 were calculated (Busing & Levy, 1957) using a theoretical absorption coefficient  $\mu=2\cdot23$  cm<sup>-1</sup>. The measurements were reduced to structure factors in the usual way (Brown & Levy, 1964). The squares of these structure factors, which have been corrected for extinction as described below, are listed in Table 1 together with their standard errors.

### Calculations

Approximate hydrogen positions were determined from a Fourier synthesis using observed structure-factor magnitudes with signs based on the heavy-atom coordinates reported by Sundera-Rao, Turley & Pepinsky (1957).

In the early stages of least-squares refinement it became apparent that the data suffer from the effects of

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severe extinction. Extinction corrections were therefore included in the refinement using the expression of Zachariasen (1967) in anisotropic form as described by Coppens & Hamilton (1970). The assumption of Type II extinction (dominated by mosaic block size)

gave a slightly better fit than did Type I (dominated by angular mosaic spread).

The calculated extinction factors by which the squares of the observed structure factors have been divided are included in Table 1. Apparently extinction has

Table 1. Observed squares of the structure factors for urea-phosphoric acid

The columns headed H, OBS, SIG, and EXT contain, respectively, the index h, the observed F<sup>2</sup> (corrected for extinction) and its standard error in units of 10<sup>-25</sup> cm<sup>2</sup>, and the calculated extinction factor (x 10<sup>3</sup>). The original uncorrected value of F<sup>2</sup> and its standard error may be obtained by multiplying OBS and SIG by EXT. The indices k and l are given in the subheadings.

Table with multiple columns for H, OBS, SIG, EXT and subheadings for k and l. The table contains a large grid of numerical data points.

reduced the intensities of 382 of the 1926 measured reflections to less than half their ideal values. The 72 strong reflections with calculated extinction factors less than 0.289 were omitted from the refinement, and these are identified by the letter *E* in Table 1.

The 155 parameters adjusted in the final cycle of refinement include the scale factor, six anisotropic extinction coefficients, the coordinates and anisotropic temperature factor coefficients for each of 16 atoms, and the neutron-scattering factors for four kinds of atoms relative to that for oxygen which was held con-

stant. The values of these parameters are listed in Table 2 together with their standard errors. The discrepancy factor based on  $F^2$ ,  $R = \sum |F_{\text{obs}}^2 - F_{\text{calc}}^2| / \sum F_{\text{obs}}^2$ , was 0.092 for the observations used in the refinement and 0.114 with all observations included.

Interatomic distances and angles are listed in Tables 3, 4, and 5, with standard errors calculated from the full variance-covariance matrix of the least-squares refinement. Table 3 also includes values corrected for the effects of thermal motion for those distances for which the riding model (Busing & Levy, 1964) seems

Table 2. Final parameters for urea-phosphoric acid

The coordinates *x*, *y*, and *z* and the temperature factor coefficients  $B_{ij}$  have been multiplied by  $10^4$  and  $10^2$ , respectively. The latter are defined by the expression for the temperature factor:

$$\exp [-(B_{11}h^2a^{*2} + \dots + 2B_{23}klb^*c^*)/4].$$

	<i>x</i>	<i>y</i>	<i>z</i>	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
P	3109 (1)	2783 (3)	3093 (2)	145 (8)	123 (9)	109 (8)	1 (6)	10 (5)	-12 (6)
O(1)	3391 (1)	907 (3)	3610 (3)	196 (8)	158 (9)	276 (9)	6 (6)	-33 (7)	20 (7)
O(2)	2775 (1)	3844 (3)	4356 (2)	200 (7)	173 (8)	104 (7)	2 (6)	14 (5)	-7 (5)
O(3)	2468 (1)	2508 (4)	1934 (2)	177 (7)	332 (11)	138 (7)	25 (7)	-42 (6)	-57 (7)
O(4)	3795 (1)	3662 (3)	2381 (2)	226 (8)	203 (9)	193 (8)	-62 (6)	78 (6)	-34 (6)
O(5)	4470 (1)	6340 (3)	3107 (2)	294 (9)	255 (10)	138 (8)	-111 (7)	19 (6)	-4 (6)
N(1)	5077 (1)	7830 (2)	4913 (2)	212 (6)	249 (8)	233 (6)	-78 (4)	4 (4)	-45 (4)
N(2)	3966 (1)	6363 (2)	5423 (2)	244 (6)	345 (9)	204 (6)	-97 (5)	81 (4)	-65 (5)
C	4500 (1)	6835 (2)	4469 (2)	159 (7)	146 (8)	143 (7)	-16 (5)	16 (5)	-14 (5)
H(1)	2977 (2)	90 (5)	3939 (5)	343 (17)	211 (15)	291 (16)	-32 (13)	-21 (13)	20 (12)
H(3)	2607 (2)	1981 (6)	951 (4)	291 (15)	300 (17)	193 (13)	20 (12)	-24 (11)	-40 (12)
H(4)	4088 (3)	5048 (6)	2815 (5)	344 (18)	374 (19)	240 (15)	44 (15)	70 (14)	45 (13)
H(11)	5497 (3)	8080 (7)	4193 (6)	282 (16)	390 (21)	361 (20)	-101 (14)	52 (15)	-22 (16)
H(12)	5120 (3)	8219 (7)	5973 (6)	404 (21)	430 (23)	306 (20)	110 (17)	-59 (16)	-106 (17)
H(21)	3518 (3)	5643 (8)	5092 (5)	307 (18)	482 (26)	335 (19)	-132 (18)	62 (15)	-61 (17)
H(22)	3987 (3)	6793 (9)	6475 (5)	470 (24)	613 (30)	242 (18)	-128 (21)	104 (16)	-129 (18)

Anisotropic extinction parameters†

$W''_{11}$	$W''_{22}$	$W''_{33}$	$W''_{12}$	$W''_{13}$	$W''_{23}$
17 (2)	9 (1)	19 (3)	7 (1)	2 (1)	3 (1)

Scattering amplitudes, ‡ cm × 10<sup>-15</sup>

P	O	N	C	H
507 (7)	577	926 (6)	658 (7)	-370 (4)
[510]	[577]	[940]	[665]	[-372]

† These are the values of  $W'_{ij}$  (Coppens & Hamilton, 1970) multiplied by  $10^4$ .

‡ Amplitudes of atoms of the same chemical element were constrained to be equal. That of oxygen was not varied. The figures in brackets are the values tabulated by the Neutron Diffraction Commission (1969).

Table 3. Interatomic distances (Å) for bonded atoms and for atoms involved in hydrogen bonds

Corrected distances are the estimated mean distances assuming the riding model for thermal motion.

Phosphoric acid	Uncorrected distance	Corrected distance	Urea	Uncorrected distance	Corrected distance
P—O(1)	1.560 (3)		C—O(5)	1.281 (3)	
P—O(2)	1.507 (3)		C—N(1)	1.323 (2)	
P—O(3)	1.553 (3)		C—N(2)	1.324 (2)	
P—O(4)	1.519 (3)		N(1)—H(11)	1.002 (5)	1.021 (5)
O(1)—H(1)	0.997 (5)	1.007 (5)	N(1)—H(12)	1.001 (5)	1.026 (5)
O(3)—H(3)	0.999 (4)	1.005 (5)	N(2)—H(21)	1.003 (5)	1.020 (5)
O(2)···H(1)'	1.664 (5)		N(2)—H(22)	1.000 (5)	1.032 (6)
O(2)···H(3)'	1.591 (4)		O(5)···H(4)	1.207 (6)	
O(2)···H(21)	1.992 (5)		O(5)···H(12)'	2.262 (6)	
O(4)···H(4)	1.223 (6)		O(5)···H(22)'	2.200 (6)	
O(4)···H(11)'	1.939 (5)				

appropriate. Table 6 lists the displacements of atoms of the urea moiety and of some surrounding atoms from the plane defined by N(1), N(2), and O(5). The standard errors of these quantities were calculated using the full variance-covariance matrix.

Fig. 1 is a stereoscopic drawing showing the packing of the molecules in a unit cell and the hydrogen bonds between them. Fig. 2, also a stereoscopic drawing, shows the basic asymmetric unit. The atoms are depicted as 50% probability ellipsoids (Johnson, 1965)

Table 4. Bond angles and angles around atoms accepting hydrogen bonds ( $^{\circ}$ )

Phosphoric acid		Urea	
O(1)—P—O(2)	111.9 (2)	O(5)—C—N(1)	118.9 (2)
O(1)—P—O(3)	108.2 (2)	O(5)—C—N(2)	121.0 (2)
O(1)—P—O(4)	105.2 (2)	N(1)—C—N(2)	120.1 (2)
O(2)—P—O(3)	106.9 (2)	C—N(1)—H(11)	118.6 (3)
O(2)—P—O(4)	113.7 (2)	C—N(1)—H(12)	120.7 (3)
O(3)—P—O(4)	110.8 (2)	H(11)—N(1)—H(12)	120.4 (4)
P—O(1)—H(1)	114.0 (3)	C—N(2)—H(21)	120.8 (3)
P—O(3)—H(3)	117.9 (3)	C—N(2)—H(22)	120.2 (3)
P—O(2)—H(1)'	115.9 (2)	H(21)—N(2)—H(22)	118.9 (4)
P—O(2)—H(3)'	123.3 (2)	C—O(5)···H(4)	117.6 (3)
P—O(2)···H(21)	110.4 (2)	C—O(5)···H(12)'	138.5 (2)
H(1)' <sup>+</sup> —O(2)···H(3)'	105.8 (2)	C—O(5)···H(22)'	118.2 (2)
H(1)' <sup>+</sup> ···O(2)···H(21)	102.8 (2)	H(4)···O(5)···H(12)'	102.4 (3)
H(3)' <sup>+</sup> ···O(2)···H(21)	94.9 (2)	H(4)···O(5)···H(22)'	98.4 (3)
P—O(4)···H(4)	124.7 (3)	H(12)···O(5)···H(22)'	62.4 (2)
P—O(4)···H(11)'	136.4 (2)		
H(4)···O(4)···H(11)'	93.7 (3)		

Table 5. Hydrogen bonds. Uncorrected interatomic distances ( $\text{\AA}$ ) and angles ( $^{\circ}$ ) at the hydrogen atom

X—H···O	X···O	X—H	H···O	$\angle$ at H
O(1)—H(1)···O(2)'	2.658 (3)	0.997 (5)	1.664 (5)	173.7 (4)
O(3)—H(3)···O(2)'	2.589 (3)	0.999 (4)	1.591 (4)	176.4 (4)
O(4)···H(4)···O(5)	2.421 (3)	1.223 (6)	1.207 (6)	169.9 (4)
N(1)—H(11)···O(4)'	2.935 (3)	1.002 (5)	1.939 (5)	172.4 (5)
N(1)—H(12)···O(5)'	3.131 (3)	1.001 (5)	2.262 (6)	144.4 (4)
N(2)—H(21)···O(2)	2.982 (2)	1.003 (5)	1.992 (5)	168.9 (5)
N(2)—H(22)···O(5)'	3.095 (3)	1.000 (5)	2.200 (6)	148.3 (5)

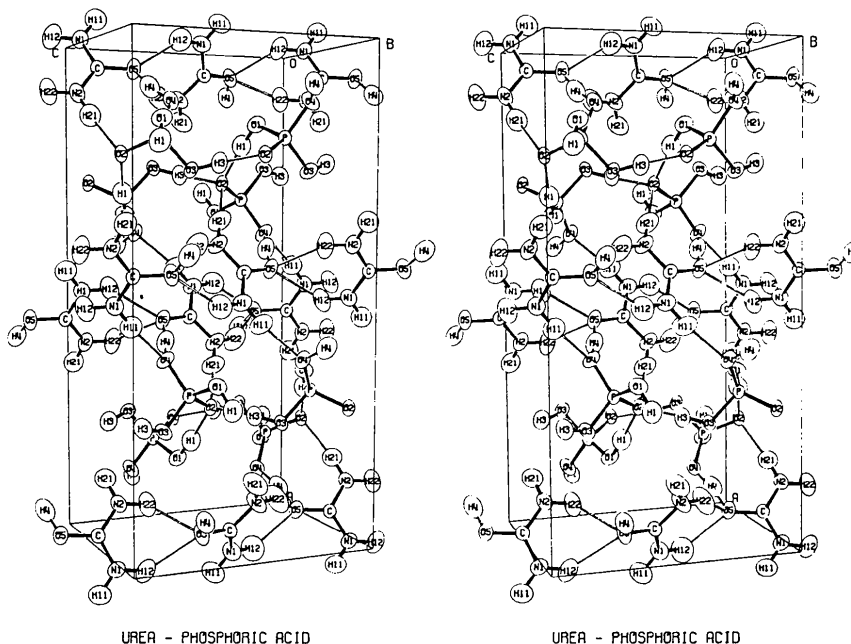


Fig. 1. Stereoscopic view showing the arrangement of the molecules in a unit cell and the hydrogen bonds between them.

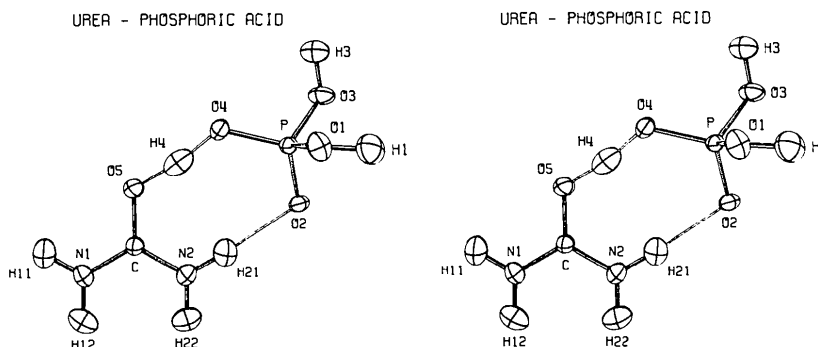


Fig. 2. Stereoscopic view of the basic asymmetric unit. The atoms are represented as 50% probability ellipsoids.

Table 6. *Displacements (Å) of atoms of the urea moiety and of some surrounding atoms from the plane defined by N(1), N(2), and O(5)*

Opposite the hydrogen atoms are listed the oxygen atoms which are acceptors for the corresponding hydrogen bonds. The displacement of atom C from this plane is 0.000 (2) Å.

Hydrogen atom		Acceptor atom	
H(4)	0.392 (6) Å	O(4)	0.894 (6) Å
H(11)	0.080 (6)	O(4')	0.044 (5)
H(12)	0.016 (6)	O(5')	-0.414 (5)
H(21)	-0.030 (6)	O(2)	0.256 (5)
H(22)	-0.032 (7)	O(5')	-0.414 (5)

with ellipses showing the principal planes of thermal motion.

Computer programs used for these calculations included the following (see *World List of Crystallographic Computer Programs*, 1972);

least-squares refinement	<i>ORXFLS3</i> ,
distances and angles	<i>ORFFE3</i> ,
structure drawing	<i>ORTEP2</i> .

Table 1 was prepared using program *EDIT* by G. M. Brown of this Laboratory.

### Discussion

The most notable feature of the structure is the short hydrogen bond which links the urea and phosphoric acid moieties. The O(4)···O(5) separation for this bond is 2.421 Å (Table 5), and the acid proton H(4) is essentially centered at distances of 1.223 Å and 1.207 Å, respectively, from these atoms. The thermal motion of H(4), which is shown in Fig. 2, has its largest principal component within 11° of the O(4)···O(5) direction. The root-mean-square displacement in this direction is 0.23 Å.

The phosphate group is essentially tetrahedral, with O-P-O angles ranging from 105.2° to 113.7°, as shown in Table 4. The distances P-O(1) and P-O(3) involving the hydroxyl oxygen atoms are 1.560 Å and 1.553 Å, respectively (Table 3). These are considerably longer than the 1.507 Å length of P-O(2), which has more

multiple-bond character because O(2) has no hydrogen atom of its own. Intermediate between these extremes is the P-O(4) distance of 1.519 Å, which seems reasonable because O(4) shares its hydrogen atom with O(5) of the urea molecule.

Worsham & Busing (1969) showed that the C-O and C-N distances in the uronium ion should be longer and shorter, respectively, than the corresponding distances in urea. The bond lengths found in the present study are, as would be expected, intermediate between those of urea (Pryor & Sanger, 1970) and those of uronium nitrate (Worsham & Busing, 1969). In urea, urea-phosphoric acid, and uronium nitrate, respectively, the C-O distances are 1.260, 1.281, and 1.298 Å, and the average C-N distances are 1.352, 1.324, and 1.314 Å.

The carbon atom of urea lies almost exactly in the plane defined by N(1), N(2), and O(5). Apparently the shared proton H(4) has little tendency to remain in the molecular plane, and it is displaced 0.392 Å to a position which makes the hydrogen bond nearly straight (Table 6). The deviations of the other urea hydrogen atoms from the plane are considerably smaller, and they do not seem to be related simply to the positions of the oxygen atoms to which they form hydrogen bonds.

Each hydrogen atom is involved in one hydrogen bond as described in Table 5, and the over-all bonding scheme is shown in Fig. 1. It is interesting to note that atoms O(1) and O(3), which have their own hydrogen atoms, accept no hydrogen bonds, but atom O(2), which has no hydrogen of its own, accepts three such bonds. Intermediate between these extremes is the atom O(4), which shares H(4) and accepts one other hydrogen bond, and the atom O(5), which also shares H(4) and accepts two rather long and bent hydrogen bonds.

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## An X-ray Structure Redetermination of the Urea-Phosphoric Acid Compound\*

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The crystal structure of the urea-phosphoric acid compound  $\text{CO}(\text{NH}_2)_2 \cdot \text{H}_3\text{PO}_4$  has been redetermined (*Pbca*;  $z=8$ ;  $a=17.6508$ ,  $b=7.4817$ ,  $c=8.9991$  Å). With a very short and nearly centred hydrogen bond from the acid to the oxygen atom of the base the structure is between that of a true adduct and that of an onium salt. All hydrogen atoms were located; with 1126 independent Cu  $K\alpha$  counter data the final  $R$  value is 0.031.

As part of our program to study the structural behaviour of the acid proton in crystalline acid/base systems (*e.g.* Altenburg & Mootz, 1971; Mootz & Altenburg, 1971) we have carried out an accurate crystal structure determination of the urea-phosphoric acid compound. Earlier investigations by Sundera-Rao, Turley & Pepinsky (1957) and by Wolfram, Arutunian, Antishkina & Porai-Koshits (1967) revealed a very strong hydrogen bond between the acid and the oxygen atom of the urea moiety ( $\text{O} \cdots \text{O}$  distance 2.414 Å and 2.45 Å respectively), but did not decide between nor comment on the structural alternatives of a true adduct and an onium salt. In order to study this aspect of the structure in detail, which requires a knowledge of the hydrogen atom positions as well as more accurate interatomic distances and angles of the heavy atoms, we decided to collect new X-ray data and repeat the analysis.

After we had finished the refinement we learned from the title programme of a meeting of a recent neutron diffraction study of the same compound by Busing & Kostansek (1971). This work has now been published in detail (Kostansek & Busing, 1972). Apart from being more accurate in hydrogen atom positions and thermal parameters, it is in good agreement with

our own study, the presentation of which can therefore be kept short.

### Experimental and crystallographic data

The urea-phosphoric acid compound crystallizes in space group *Pbca* with eight formula units  $\text{CO}(\text{NH}_2)_2 \cdot \text{H}_3\text{PO}_4$  per unit cell. A least-squares refinement of angular measurements on a diffractometer yielded the improved lattice parameters  $a=17.6508$  (9),  $b=7.4817$  (4) and  $c=8.9991$  (7) Å. The intensities of 1126 independent reflexions with  $\theta < 70^\circ$  were measured with Ni-filtered Cu  $K\alpha$  radiation on a tape-controlled Siemens automatic diffractometer in the  $\theta/2\theta$  mode. No correction for absorption was applied ( $\mu_{\text{Cu } K\alpha} = 38.9 \text{ cm}^{-1}$ ).

### Refinement of the structure

Least-squares refinement was started from the heavy atom parameters of Sundera-Rao, Turley & Pepinsky (1957). The seven hydrogen atoms could be located in a difference Fourier map. After an empirical correction of the thirteen strongest intensities for extinction, anisotropic treatment of all atoms converged to an unweighted  $R$  value of 0.031. The observations were weighted according to  $w=1$  for  $|F_o| < K$  and  $w=K^2/|F_o|^2$  for  $|F_o| \geq K$  with  $K=20$ . The form factors used

\* Short version of a part of the dissertation of K.-R. Albrand, Techn. Universität Braunschweig, (1972).